

Neutral Atoms - Practice

Name: _____

Date: _____

Hour: _____

Directions:

Fill in the missing information in the chart below.

Element Symbol	Atomic #	Atomic Mass	Protons	Neutrons	Electrons	Element Symbol	Atomic #	Atomic Mass	Protons	Neutrons	Electrons
H							4				4
Li							6	12		6	6
F							7		7		7
Ne							16				16
Mg							43				43
K						Ca					
Sc							87	223	87	136	87
Co							92				
Ni							14	28	14	14	14
Ga											88
Br							104	261	104	157	104
Kr											94
Y							23	51	23	28	23
Mo						Ba					
Ag							56				56
Sn							40				40
I							9	19	9	10	9
Xe						Cl					

Isotopes and Ions – Practice

Name: _____

Date: _____

Hour: _____

Determine if the elements in the pair are “isotopes” or “different elements.”

1. Element A has 2 protons and 3 neutrons. _____
Element B has 3 protons and 2 neutrons.
2. Element C has 2 protons and 3 neutrons. _____
Element D has 2 protons and 4 neutrons.
3. Element E has 71 protons and 71 neutrons. _____
Element F has 71 protons and 68 neutrons.
4. Element G has 65 protons and 61 neutrons. _____
Element H has 65 neutrons and 61 protons.
5. Element H has an atomic # of 15 and a mass of 30. _____
Element J has an atomic # of 15 and a mass of 31.
6. Element K has 8 protons and 8 neutrons. _____
Element L has an atomic # 8 and 8 neutrons.
7. Element M has an atomic # of 84 and a mass of 138. _____
Element N has 84 protons and 43 neutrons.

Conclusion:

In your own words, explain how you can tell if two atoms are isotopes or different elements.

Directions:

Fill in the missing information below.

Atom	Protons	Electrons	Charge
O	2	2	0
P	2		-1
Q	9		-4
R	12		+3
S	20		+2

Atom	Protons	Electrons	Charge
T	14	12	
U	98	89	
V	26	25	
W		45	-4
X		50	+6

Conclusion:

In your own words, explain how you determine the charge of an atom.