**Chapter 10 Review: Quiz is Friday, June 2**

1. Determine oxidation numbers for the following:

(a) CaF2 (b) Ag3P (c) Sc2S3 (d) CO

(e) CO32-  (f) H2O (g) NaH (h) NH4+

(i) BaSO4 (j) Cu(MnO4)2 (k) Cr(OH)4  (l) KCN

2. Which of the following is a redox reaction? Explain how you determined this?

a) CaCO3 → CaO + CO2

b) NaOH + HCl → NaCl + H2O

c) 2NH4Cl + Ca(OH)2 → 2 NH3 + 2 H2O + CaCl2

d) 2 H2O → 2 H2 + O2

3. In the reaction Mg+ Cl2 → MgCl2, the correct half-reaction for the oxidation that occurs is:

A. Mg + 2e- → Mg2+ B. Mg → Mg2++ 2e- C. Cl2+ 2e- → 2 Cl- D. Cl2→ 2 Cl- + 2e-

4. Which half-reaction correctly represents a reduction reaction?

1. Sn0 + 2e- → Sn2+ B. Na0 + e- → Na1+ C. Li0 + e- → Li+ D. Br20 + 2e- → 2Br –

5. Which half-cell reaction correctly represents oxidation?

1. Pb2+ + 2e- → Pb B. Pb + 2e- → Pb2+ C. Pb2+ → Pb + 2e- D. Pb→ Pb2+ + 2e-

6. Where does oxidation occur?

7. Where does reduction occur?